Date Planned : __ / __ / __	Daily Tutorial Sheet - 7	Expected Duration : 90 Min
Actual Date of Attempt : __ / __ / __	Level - 2	Exact Duration :_____

86. For a given principal level n = 5, the energy of its subshells is of the order:
(A) s < d < f < p **(B)** s < p < d < f **(C)** d < f < p < s **(D)** s < p < f < d

87. The maximum number of electrons in an orbital, is governed by.....principle.
(A) Hund's **(B)** Aufbau **(C)** Pauli's **(D)** None of these

88. Any p orbital can accommodate up to:
(A) Four electrons **(B)** Two electrons with parallel spin
(C) Six electrons **(D)** Two electrons with opposite spin

89. The spectral line obtained when an electron jumps from n = 6 to n = 2 level in hydrogen atom belongs to the:
(A) Balmer series **(B)** Lyman series **(C)** Paschen series **(D)** Pfund series

90. The ionization potential of hydrogen atom is 13.6 eV. The energy required to remove an electron from the n = 2 state of the hydrogen atom is:
(A) 3.4 eV **(B)** 6.8 eV **(C)** 13.6 eV **(D)** 27.2 eV

91. The e/m ratio for electron was determined by.....
(A) Einstein **(B)** Newton **(C)** J J. Thomson **(D)** Planck

92. $Na^{\oplus}$ and Ne are ....to each other:
(A) Isobar **(B)** Isoelectronic **(C)** Isotone **(D)** Isotope

93. Which of the following angular momentum is not possible for electron in Bohr's orbit?
(A) $0.5\hbar$ **(B)** $\hbar$ **(C)** $2\hbar$ **(D)** $3\hbar$

94. If the radius of the first Bohr orbit of the H atom is r then for the Li^{2+} ion it will be:
(A) 3r **(B)** 9r **(C)** r/3 **(D)** r/9

95. The wavelength of certain line in H-atom spectra is observed to be 4341Å. ($R_H = 109677 cm^{-1}$). The value of quantum number of higher state is:
(A) 3 **(B)** 4 **(C)** 5 **(D)** Data insufficient